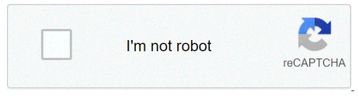

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What are all the metalloids on the periodic table

Metalloids are elements from the periodic table with properties that lie between metals and non-metals. The following ScienceStruck article will cover some information related to metalloids. Periodic table was based on the atomic weight of the elements. He found that there were many elements that shared similar properties and occurred periodically. Thus, he came up with the concept of periodic table of elements. His table had been divided into three main sections – metals, nonmetals, and metalloids. In this article, we shall concentrate on the metalloids from the periodic table. According to Mendeleev's law of periodic table, the chemical and physical properties of elements vary in a periodic manner according to their atomic weights. However, the modern periodic table of elements follow the law that, the properties of elements are a periodic function of their atomic numbers. The chemical elements of the same group had similar properties. There are different regions in the periodic table that are called periodic table blocks, as they are named after the name of the atom. Location of Metalloids in the Periodic Table. The metalloids, also known as semi-metals, are placed between metals and non-metals in the periodic table of elements. There are seven elements that are classified as metalloids and placed in Group 13, 14, 15, 16, and 17. They are found in a stair step line that helps differentiate metals from non-metals in this element table. The line that separates metalloids from the metals and non-metals in the periodic table is called amphoteric line. List of Metalloid Elements As mentioned above, there are 7 elements in the periodic table that exhibit metalloid properties. They occur in a diagonal line from boron to astatine through the p-block. The elements in the upper-right portion of the line show increasing non-metallic behavior and the elements at the lower-left of the line show increasing metallic behavior. The list of metalloids in the periodic table are as follows: Boron (B) Silicon (Si) Germanium (Ge) Arsenic (As) Antimony (Sb) Tellurium (Te) Polonium (Po) Metalloids in the Periodic Table s-block Transition Elementsd- block p-block NobelElement Group 1 2 3 4 5 6 7 8 9 10 11 12 13 14 15 16 17 18 1 1H1.0079 2He4.0026 2 3Li6.941 4Be9.0122 5B10.811 6C12.011 7N14.007 8O15.999 9F18.998 10Ne20.180 3 11Na22.990 12Mg24.305 13Al26.982 14Si28.086 15P30.974 16S32.066 17Cl35.453 18Ar39.948 4 19K39.098 20Ca40.078 21Sc44.956 22Ti47.867 23V50.942 24Cr51.996 25Mn54.938 26Fe55.845 27Co58.933 28Ni58.693 29Cu63.546 30Zn65.409 31Ga69.723 32Ge72.64 33As74.922 34Se78.96 35Br79.904 36Kr83.798 5 37Rb85.468 38Sr87.62 39Y88.906 40Zr91.224 41Nb92.906 42Mo95.94 43Tc(98) 44Ru101.07 45Rh102.91 46Pd106.42 47Ag107.87 48Cd112.41 49In114.82 50Sn118.71 51Sb121.76 52Te127.60 53I126.90 54Xe131.29 6 55Cs132.91 56Ba137.33 57 – 71La-Lu 72Hf178.49 73Ta180.95 74W183.84 75Re186.21 76Os190.23 77Ir192.22 78Pt195.08 79Au196.97 80Hg200.59 81Tl204.38 82Pb207.2 83Bi208.98 84Po(209) 85At(210) 86Rn(222) 7 87Fr(223) 88Ra(226) 89 -103Ac-Lr 104Rf(261) 105Db(262) 106Sg(266) 107Bh(264) 108Hs(277) 109Mt(268) 110Ds(281) 111Rg(272) 112Cn(277) Lanthanide 57La138.91 58Ce140.12 59Pr140.91 60Nd144.24 61Pm(145) 62Sm150.36 63Eu151.96 64Gd157.25 65Tb158.93 66Dy162.50 67Ho164.93 68Er167.26 69Tm168.93 70Yb173.04 71Lu174.97 Actinide 89Ac(227) 90Th232.04 91Pa(231)92U(238)93Np(237)94Pu(244)95Am(243)96Cm(247)97Bk(247) 98Cf(251)99Es(252)100Fm(257)101Md(258)102No(259)103Lr(262) Alkali Metals Alkaline Earth Metals Lanthinides Actinides Transition Metals Poor Metals Other Metals Nobel Gases Metalloids Unknown Chemical Properties The term metalloid comes from the Greek word metallon, which means 'metal', and edios, meaning 'sort.' The metalloids are often seen forming amphoteric oxides, and they behave as semiconductors. They have properties of both metals and non-metals in the periodic table. They even carry electric charge that makes them suitable for use in computers and calculators. Their ionization energy as well as electronegativity values are between those of metals and non-metals. Their reactivity depends on the metals that they are reacting with. As you can see with the first version of periodic table in 1864. Mendeleev's